

Periodic Table Quiz

1. Early chemists used the planets to identify the elements known to them. This later was a problem, when more elements were discovered, because they ran out of planets.

This symbol  represent the planet and element ...

- A. Mars - iron
 B. Venus - copper
 C. Mercury - mercury
 D. Jupiter - tin

2.

Dalton's 1808AD symbols and formulae.

| | | |
|--|---|--|
|  Hydrogen |  Soda |  Ammonia |
|  Nitrogen |  Pot Ash |  Olefiant |
|  Carbon |  Oxygen |  Carbonic Oxide |
|  Sulphur |  Copper |  Carbonic Acid |
|  Phosphorus |  Lead |  Sulphuric Acid |
|  Alumina |  Water | |

The purpose for developing a new set of chemical symbols was that Dalton wanted to ...

- A. become rich and famous
 B. win a Nobel Prize in Chemistry
 C. earn bragging rights with other scientists
 D. streamline communication with other scientists
3. Berzelius later revised Dalton's symbols by replacing the pictures with ...
- A. letters
 B. names
 C. shapes
 D. numbers
4. John Newland's "*law of octaves*" identified the pattern in which the properties of the elements seemed to repeat at regular intervals, they were similar to the ...
- A. base ten number system
 B. suits of playing cards
 C. heartbeats of different animals
 D. octave scale in music
5. These elements have both metal and non-metal properties. Some of them are semi-conductors, which means, they can carry an electrical charge under special conditions. Making them great for computers and calculators. They are the ...
- A. Rare Earth Elements
 B. Transition Metals
 C. Other Metals
 D. Metalloids
6. The 6 elements in this group all have the maximum number of electrons possible in their outer shell which makes them stable. They are known as the ...
- A. Halogens
 B. Alkali Metals
 C. Noble Gases
 D. Alkaline Earth Metals
7. Dmitri Mendeleev wanted to find a pattern that would allow him to predict the properties of elements not yet discovered. By using information cards he charted the pattern that seemed to work. The characteristic that showed that the properties of elements vary periodically was the ...
- A. atomic number
 B. atomic mass
 C. symbol
 D. density

Modern Periodic Table
Excerpt for Students

| | | | | | | | |
|--|---|--|--|---|---|--|--|
| 1 1.01 1+ H Hydrogen | | | | | | | 2 4.00 He Helium |
| 3 6.94 1+ Li Lithium | 4 9.01 2+ Be Beryllium | 5 10.81 ~ B Boron | 6 12.01 ~ C Carbon | 7 14.01 3- N Nitrogen | 8 16.00 2- O Oxygen | 9 19.00 1- F Fluorine | 10 20.18 ~ Ne Neon |
| 11 22.99 1+ Na Sodium | 12 24.31 2+ Mg Magnesium | 13 26.98 3+ Al Aluminum | 14 28.09 ~ Si Silicon | 15 30.97 3- P Phosphorus | 16 32.07 2- S Sulfur | 17 35.45 1- Cl Chlorine | 18 39.95 ~ Ar Argon |

| | |
|--------------------|------------------------|
| Atomic Number → 14 | 28.09 ← Atomic Mass |
| Symbol → Si | ~ ← ion charge |
| | Silicon ← Element Name |

8. In an element, the number of protons in the nucleus determines the ...
 - A. **atomic number**
 - B. **atomic mass**
 - C. **symbol**
 - D. **density**

9. Each letter, or combination of letters in the Periodic table represent the atomic ...
 - A. **number**
 - B. **mass**
 - C. **symbol**
 - D. **name**

10. Periods are elements in the periodic table that can be found in the same ...
 - A. **horizontal columns**
 - B. **horizontal rows**
 - C. **vertical columns**
 - D. **vertical rows**

11. Alkali metals are found in group 1 in the Periodic table. One characteristic of these elements is that they are ...
 - A. **stable**
 - B. **highly reactive**
 - C. **magnetic**
 - D. **radioactive**

12. The 38 Transition metals are the only elements known to produce a ...
 - A. **strong reaction**
 - B. **magnetic field**
 - C. **stable compound**
 - D. **electric charge**

13. The elements in this group - because they have metal and non-metal properties - as semi-conductors they are used in ...
 - A. **computers**
 - B. **blast furnaces**
 - C. **making salts**
 - D. **synthetics**

14. The mass number of each element is determined by the combination of these in the nucleus ...
 - A. **protons and electrons**
 - B. **electrons and neutrons**
 - C. **neutrons and protons**
 - D. **protons, neutrons and elements**

15. The stability of an element is determined by the number of electrons the element has in its ...
 - A. **inner shell**
 - B. **outer shell**
 - C. **nucleus**
 - D. **atomic number**

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